

1. Isotopes of boron - B-10 B-11 \rightarrow % of abundance -20% and 80% Find AAM of B.
2. If an element M exists in two isotopic forms $50M$ & $52M$. The average atomic mass of isotopic mixture is 51.7. Calculate abundance of each isotopic forms.
3. Helium has two naturally occurring isotopes, He-3 and He-4. The atomic mass of helium is 3.98 amu. Which isotope is more abundant in nature?